

COURT OF THE PROPERTY OF THE P

MICROCOPY RESOLUTION TEST CHART
NATIONAL BUREAU OF STANDARDS-1963-A



OFFICE OF NAVAL RESEARCH

Contract N00014-80-C-0472

Task No. NR 056-749

TECHNICAL REPORT No. 37

Laser-Induced Bound States at Surfaces: Ion Neutralization and Adsorption

by

Kai-Shue Lam, Michael Hutchinson and Thomas F. George

Prepared for Publication

in

Collisions and Half-Collisions with Lasers edited by N. K. Rahman and C. Guidotti (Harwood Academic, Chur, Switzerland)

Department of Chemistry University of Rochester Rochester, New York 14627



June 1983

Reproduction in whole or in part is permitted for any purpose of the United States Government.

This document has been approved for public release and sale; its distribution is unlimited.

83 07 12 064

SECURITY CLASSIFICATION OF THIS PAGE (When Data Entered)

REPORT DOCUMENTATION PAGE		READ INSTRUCTIONS BEFORE COMPLETING FORM
1. REPORT NUMBER		RECIPIENT'S CATALOG NUMBER
UROCHESTER/DC/83/TR-37	A130 27	
4. TITLE (and Subility) Laser-Induced Bound States at Surf	'acas :	5. TYPE OF REPORT & PERIOD COVERED
Ion Neutralization and Adsorption	aces.	
	!	6. PERFORMING ORG. REPORT NUMBER
7. Author(*) Kai-Shue Lam, Michael Hutchinson a		8. CONTRACT OR GRANT NUMBER(s)
Thomas F. George	na ,	N00014-80-C-0472
9. PERFORMING ORGANIZATION NAME AND ADDRESS Department of Chemistry	!	10. PROGRAM ELEMENT, PROJECT, TASK AREA & WORK UNIT NUMBERS
University of Rochester	!	NR 056-749
Rochester, New York 14627		
Office of Naval Research	!	June 1983
Chemistry Program Code 472 Arlington, Virginia 22217	!	13. NUMBER OF PAGES
AFTINGLON, VIFGINIA 22217 14. MONITORING AGENCY NAME & ADDRESS(If differen	it from Controlling Office)	15. SECURITY CLASS. (of this report)
	1	
	ı	Unclassified 154. DECLASSIFICATION/DOWNGRADING SCHEDULE
		SCHEDULE
16. DISTRIBUTION STATEMENT (of this Report) This document has been approved for	r nublic release	and sale: its distribution
is unlimited.	Post of Tarana	ma outer too atom.
17. DISTRIBUTION STATEMENT (of the abetract entered	in Block 20, If different fre	im Report)
·		
18. SUPPLEMENTARY NOTES		
Prepared for publication in Collis	ions and Half-Co	llisions with Lasers
edited by N. K. Rahman and C. Guide	otti (Harwood Ac	ademic, Chur, Switzerland)
19. KEY WORDS (Continue on reverse elde il necessary an LASER-INDUCED BOUND STATES	nd Identify by block number) QUANTUM ME	
SURFACES		THS AND SHIFTS
ION NEUTRALIZATION	PHONONS	
ADSORPTION		
20. ABSTRACT (Continue on reverse elde if necessary and Two mechanisms for laser-generated	d identify by block number)	f a foreign energies on a
1WO Mechanisms for laser-yeneraced	1 pontin 2 races of	i a iniciôn sheries on a

solid surface are discussed. The first, ion neutralization, leads to a bound electronic valence state of a projectile ion that is not degenerate with any electronic band states of the surface, while the second, radiative adsorption, gives a stable vibrational state of an adsorbed atom. The laser intensity plays the dominant role in the first process, whereas in the second a resonantly-tuned frequency is of greater importance.

Collisions and Half-Collisions with Lasers edited by N. K. Rahman and C. Guidotti (Harwood Academic, Chur, Switzerland)

LASER-INDUCED BOUND STATES AT SURFACES: ION NEUTRALIZATION AND ADSORPTION

KAI-SHUE LAM, MICHAEL HUTCHINSON AND THOMAS F. GEORGE Department of Chemistry, University of Rochester Rochester, New York 14627 USA

Abstract A laser can be used to generate bound states, both electronic and vibrational, of a foreign atom on a solid surface, and is capable of enhancing processes like ion neutralization and adsorption.

INTRODUCTION

Two mechanisms for laser-generated bound states of a foreign species on a solid surface are discussed. The first, ion neutralization, leads to a bound electronic valence state of a projectile ion that is not degenerate with any electronic band states of the surface, while the second, radiative adsorption, gives a stable vibrational state of an adsorbed atom. The laser intensity plays the dominant role in the first process, whereas in the second a resonantly-tuned frequency is of greater importance.

ION NEUTRALIZATION

In many theories treating neutralization of ions scattered from solid surfaces, 1,2 resonance processes play a dominant role. This kind of resonance is between a discrete state and a continuum level. Thus a valence level of the



100 Codes
100 Special

THE RESERVE OF THE PROPERTY OF

projectile ion (the discrete state) is considered to have a position-dependent energy $\epsilon_0(z)$ (see Fig. 1) which, at some region of small z, is resonant with a continuum of band levels of the electronic states in the solid surface; the position z represents some measure of the distance of the ion from the surface. The bound-continuum interaction is usually assumed to be significant only for small z, that is, when the ion is near the surface. Moreover, the strength of the interaction is assumed to be such that both the shift and width of the resonant state are small, and that the energy spectrum of the ion-surface system is the same as that of the band states of the solid surface. These limitations on the bound-continuum interaction, together with the particular nature of the model -- one discrete state embedded in one continuum--lead to the situation where true bound states of the ion-surface system play no role at all in the discription of the mechanisms leading to charge transfer. If such mechanisms require true bound states, they have to be added to the model. For example, Auger neutralization may take place if a deep-lying unoccupied level of the incoming ion is available.3

Within the constraints of the model described above and a particular physical system, the bound-continuum interaction has a fixed strength, and thus true bound states either enter the picture or not at all. With the introduction of a laser, however, the situation is changed drastically. The fact of crucial importance is that both the laser frequency and the field strength are adjustable. Thus the same model, when it is understood that the bound-continuum interaction is due to a field coupling, can not only incorporate a variable $\varepsilon_0(z)$ (variable not only with respect to z but by amounts directly related to $\widetilde{\Lambda}_\omega$) but also a variable coupling strength

(directly related to the field strength). It is precisely these variable quantities which lead to the possible existence of true bound states, even when they are precluded in the absence of the field. The laser may then be used to enhance bound state mechanisms which are either unimportant or impossible in the field-free situation. With respect to the Auger neutralization process mentioned above, the laser may literally create a "deep-lying" valence state to act as receptor of an electron from a band level. Such a state may also interact resonantly with any core levels of the solid surface that happen to be approximately degenerate with it. In what follows we give a brief discussion of the theory behind the formation of the laser-induced bound state.

The schematic picture of the energy level structure in our model is given in Fig. 1. The Hamiltonian may be written as

$$H(z) = \sum_{k} \varepsilon_{k}^{n} n_{k} + (\varepsilon_{0}(z) + \hbar \omega) n_{0} + \sum_{k} [\nabla_{k}(z) c_{0}^{\dagger} c_{k} + \text{h.c.}], (1)$$

where $V_k(z)$ is the bound-continuum interaction provided by the field coupling, and k is the band index. The emergence of possible bound states is most easily seen by focusing on the eigenvalue equation for H(z):

$$\varepsilon - (\varepsilon_0 + \hbar\omega) + g^2 \int d\varepsilon' \frac{\rho(\varepsilon') |V(\varepsilon')|^2}{\varepsilon' - \varepsilon} = 0, \qquad (2)$$

where we have introduced a coupling strength g into $\mathbf{V}_{k}(\mathbf{z})$ such that

$$V_{k}(z) = gV(\varepsilon),$$
 (3)

and the z dependence is not explicitly written on the RHS. For radiative coupling, g^2 is directly proportional to the field strength.

Under what conditions will a true bound state emerge? There will be bound states when Eq. (2) admits negative

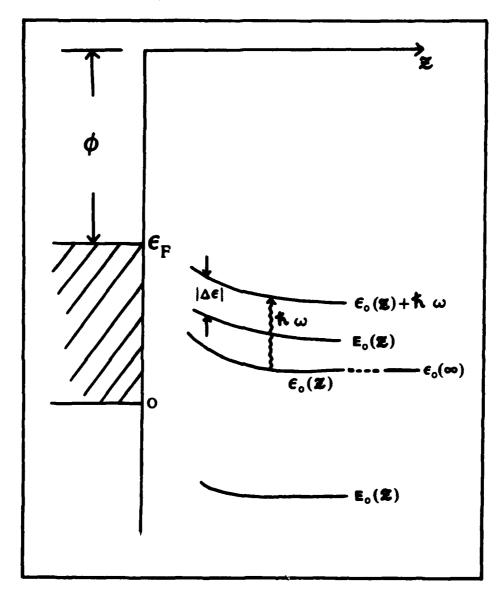


FIGURE 1. Energy level structure of the model describing ion neutralization at a solid surface. $\epsilon_0(z)$ is the unoccupied valence level of the ion, ω is the laser frequency, and $\Delta\epsilon$ is the level shift (this is shown to be negative in the picture). $E_0(z)$ is either a virtual level when it is degenerate with the conduction band, or a true bound state when it is outside the conduction band. $\epsilon_{\rm F}$ is the Fermi energy, and ϕ is the work function.

energy solutions. Let this solution be $E_0 = -x$, where x>0. Eq. (2) then reads

$$x + \varepsilon_0 + \hbar \omega = g^2 \int d\varepsilon' \frac{\rho(\varepsilon') |V(\varepsilon')|^2}{\varepsilon' + x}. \tag{4}$$

Examination of the quantitative picture for a graphical solution quickly reveals that there will be a negative energy solution only when

where

$$g_{crit} = \frac{\varepsilon_0 + 1/\omega}{\int d\varepsilon' \frac{\rho(\varepsilon') |V(\varepsilon')|^2}{\varepsilon'}}$$
(5)

Thus whenever the laser field strength is increased beyond a value specified by the critical coupling constant g_{crit} , a true bound state emerges. Furthermore, the model only admits one and only one such state.

For g < g_{crit} , however, the valence level of the projectile ion becomes an unstable (virtual) state with a shift in energy, $\Delta \varepsilon$, given by

$$\Delta \varepsilon = -g^2 P \int d\varepsilon' \frac{\rho(\varepsilon') |V(\varepsilon')|^2}{\varepsilon' - E_0}, \qquad (6)$$

where P denotes the principal value of the integral. Eq. (6) implies that when

$$|\Delta \varepsilon| > \varepsilon_0 + \kappa_0$$

we have a true bound state. Looking at Fig. 1, then, the laser can be imagined to do the following thing: As the field strength is increased, the resonant valence level ϵ_0 + $\hbar \omega$ is pulled progressively down the conduction band. As long as g < $\rm g_{crit}$, $\rm E_0(z)$ stays within the conduction band, and at most we have an unstable state. But when

 $g > g_{crit}$, $E_0(z)$ falls outside of the band, and a true bound state results.

RADIATION-ASSISTED ADSORPTION

We now consider an alternative mechanism for laser-generated bound states. This is the process of radiative adsorption, which is illustrated in Fig. 2. During the encounter with a surface, an adatom can undergo transition to a bound state of the adatom-surface potential by stimulated emission of a photon. If the surface were rigid, such a bound state would be unstable to the reverse process, photon absorption. However, by coupling the adatom motion to the phonon "bath" of the solid, there exists the possibility of a simultaneous decay to a lower-lying bound state by phonon creation. Such a state would be a true final state, except at temperatures sufficiently high that phonon annihilation (feedback) is important.

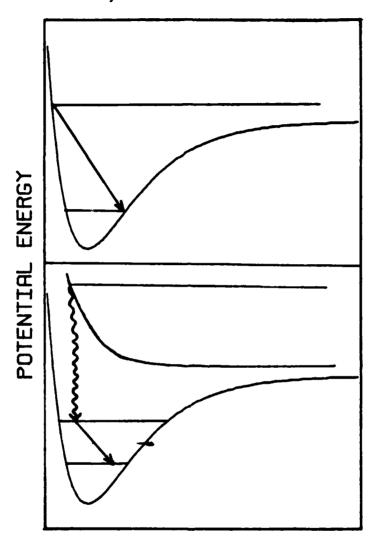
We shall now sketch the theory of this process for adsorption on a one-dimensional lattice (the extension to three dimensions is straightforward, but the notation for the one-dimensional problem is simpler). We may write the Hamiltonian as

$$H = H_0 + H_{int}^r + H_{int}^p, \tag{7}$$

where

$$H_0 = H^r + H^p + H^a,$$
 (8)

and the superscripts r, p and a stand for the radiation, phonons and the adatom, respectively. The corresponding eigenvectors of these zeroth-order Hamiltonians are written



ADATOM-SURFACE DISTANCE (z-z,)

FIGURE 2. Adsorption events depicted in the space of the atom-surface electronic eigenstates. Top: adsorption from a continuum by the creation of a single phonon. Bottom: adsorption from a continuum by the "simultaneous" creation of a photon and a phonon. The phonon events are depicted by straight arrows.

as $|n_{k\sigma}^{r}\rangle, |n_{k}^{p}\rangle$ and $|E^{+}\rangle, |f\rangle$. The third and fourth eigenvectors (of H^a) are, respectively, the scattering state at energy E (with an outgoing wave boundary condition) and a final bound state f. The latter state thus represents the adatom bound to the surface. Using a projection operator Q to project out the bound intermediate state $|m\rangle$ to which the radiative transition is made, it is now possible to write the adsorption rate R due to a plane-polarized laser as

$$R = |\langle n_{k}^{P} + 1 | \langle f | H_{int}^{P} | m \rangle | n_{k}^{P} \rangle|^{2}$$

$$\times |\langle G_{QQ}^{+} \rangle_{mm} \langle n_{k0}^{r} + 1 | \langle m | H_{int}^{r} | E^{+} \rangle | n_{k0}^{r} \rangle|^{2}$$

$$\times \delta((E - E_{f}) - \hbar \omega_{p} - \hbar \omega_{r}), \qquad (9)$$

where

$$G_{QQ}^{+} = \lim_{\epsilon \to 0} (E + i\eta - H_{QQ} - H_{QP}G_{0}^{+}H_{PQ})^{-1},$$
 (10)

$$P = 1-Q$$

and

$$O_{XY} = XOY$$
.

 G_0^+ is the Green's function for non-resonant scattering and is expanded in a product basis of surface plane waves and adatom scattering out-waves.

We now consider the contribution of a single-phonon transition to the rate by expanding the operator $H_{int\ in\ a}^{p}$ Taylor's series in the lattice coordinate z_{ρ} :

$$H_{\text{int}}^{p} = \sum_{\ell} d^{\ell}(z) (z_{\ell} - z_{\ell}^{0}),$$
 (11)

where z is the corrdinate of the as atom, and z_{ℓ}^{U} is the equilibrium (frozen lattice coordinate of atom ℓ in the

chain,

$$d^{\ell}(z) = \frac{\partial}{\partial z_{\ell}} V(z-z_{\ell}) \Big|_{z_{\ell}=z_{\ell}^{0}}. \tag{12}$$

To arrive at the total averaged rate <R>, it is necessary to average over initial and sum over final phonon states. Dropping for simplicity the Fock states of the radiation field, we are led by standard manipulations 4,5 to

$$\langle R \rangle = \sum_{\ell} \frac{\left| \langle f | \sum_{\mathbf{p}} \mathbf{d}^{\ell}(\mathbf{z}) | \mathbf{m} \rangle \right|^{2}}{NM} \frac{1}{\omega_{\mathbf{p}}} (\overline{n}_{\mathbf{p}} + 1) \rho(\omega_{\mathbf{p}})$$

$$\times \left| (G_{00}^{\dagger})_{\mathbf{mm}} \langle \mathbf{m} | H_{\mathbf{int}}^{\mathbf{r}} | E^{\dagger} \rangle \right|^{2}, \qquad (13)$$

where $\rho(\omega_p)$ is the phonon density of states as a function of the energy-conserving frequency, ω_p , $\frac{1}{p}$ is the average occupation number of phonon mode p,

$$\frac{1}{n_p} = (e^{\frac{1}{2}\pi\omega_p/kT}-1)^{-1},$$

and there are N atoms in the chain, each of mass M. An important feature of Eq. (13) is the natural separation of the gas-phase problem from that of the surface. The only point of contact is in $(G_{QQ}^+)_{mm}$, which would contain width and level shift terms due to the phonons, and related terms which are due to the radiative interaction.

ACKNOWLEDGMENTS

This research was supported by the Air Force Office of Scientific Research (AFSC), United States Air Force, under Grant AFOSR-82-0046, the Office of Naval Research and the U.S. Army Research Office. The United States Government is authorized to reproduce and distribute reprints for governmental purposes notwithstanding any copyright notation hereon. TFG acknowledges the Camille and Henry Dreyfus

Foundation for a Teacher-Scholar Award (1975-84) and the John Simon Guggenheim Memorial Foundation for a Fellowship (1983-84).

REFERENCES

- 1. J. C. Tully, Phys. Rev. B, 16, 4324 (1977).
- 2. R. Brako and D. M. Newns, Surface Sci., 108, 253 (1981).
- 3. H. D. Hagstrum, Phys. Rev., 96, 336 (1954); 122, 83 (1980).
- F. O. Goodman, <u>Prog. Surf. Sci.</u>, <u>5</u>, 261 (1974)
 A. A. Maradudin, E. W. Montroll, G. H. Weiss and I. P. Ipatova, in Theory of Lattice Dynamics in the Harmonic Approximation (Academic Press, New York, 1971), pp. 261-375.

TECHNICAL REPORT DISTRIBUTION LIST, GEN

	No.		No.
	Copies		Copies
	<u>007200</u>		<u>oopies</u>
Office of Naval Research		Naval Ocean Systems Center	
Attn: Code 413		Attn: Mr. Joe McCartney	
800 North Quincy Street		San Diego, California 92152	1
Arlington, Virginia 22217	2		
-		Naval Weapons Center	
ONR Pasadena Detachment		Attn: Dr. A. B. Amster,	
Attn: Dr. R. J. Marcus		Chemistry Division	
1030 East Green Street		China Lake, California 93555	1
Pasadena, California 91106	1		
		Naval Civil Engineering Laboratory	
Commander, Naval Air Systems Command		Attn: Dr. R. W. Drisko	
Attn: Code 310C (H. Rosenwasser)		Port Hueneme, California 93401	1
Department of the Navy			
Washington, D.C. 20360	1	Dean William Tolles	
		Naval Postgraduate School	
Defense Technical Information Center		Monterey, California 93940	1
Building 5, Cameron Station			
Alexandria, Virginia 22314	12	Scientific Advisor	
		Commandant of the Marine Corps	
Dr. Fred Saalfeld		(Code RD-1)	
Chemistry Division, Code 6100		Washington, D.C. 20380	1
Naval Research Laboratory			
Washington, D.C. 20375	1	Naval Ship Research and Development	
		Center	
U.S. Army Research Office		Attn: Dr. G. Bosmajian, Applied	
Attn: CRD-AA-IP		Chemistry Division	
P. O. Box 12211		Annapolis, Maryland 21401	1
Research Triangle Park, N.C. 27709	1		
,		Mr. John Boyle	
Mr. Vincent Schaper		Materials Branch	
DTNSRDC Code 2803		Naval Ship Engineering Center	
Annapolis, Maryland 21402	1	Philadelphia, Pennsylvania 19112	1
Naval Ocean Systems Center		Mr. A. M. Anzalone	
Attn: Dr. S. Yamamoto		Administrative Librarian	
Marine Sciences Division		PLASTEC/ARRADCOM	
San Diego, California 91232	1	Bldg 3401	_
B - B - 11 # - 11-1		Dover, New Jersey 07801	1
Dr. David L. Nelson			
Chemistry Program			
Office of Naval Research			
800 North Quincy Street	•		
Arlington, Virginia 22217	1		

TECHNICAL REPORT DISTRIBUTION LIST, 056

	No. Copies		No. Copies
Dr. G. A. Somorjai		Dr. W. Kohn	
Department of Chemistry		Department of Physics	
University of California		University of California	
Berkeley, California 94720	1	(San Diego)	
•		La Jolla, California 92037	1
Dr. J. Murday			
Naval Research Laboratory		Dr. R. L. Park	
Surface Chemistry Division (6170)		Director, Center of Materials	
455 Overlook Avenue, S.W.		Research	
Washington, D.C. 20375	1	University of Maryland	
		College Park, Maryland 20742	1
Dr. J. B. Hudson			
Materials Division		Dr. W. T. Peria	
Rensselaer Polytechnic Institute		Electrical Engineering Department	
Troy, New York 12181	1	University of Minnesota	
		Minneapolis, Minnesota 55455	1
Dr. Theodore E. Madey			
Surface Chemistry Section		Dr. Chia-wei Woo	
Department of Commerce		Department of Physics	
National Bureau of Standards		Northwestern University	_
Washington, D.C. 20234	1	Evanston, Illinois 60201	1
Dr. J. M. White		Dr. Robert M. Hexter	
Department of Chemistry		Department of Chemistry	
University of Texas		University of Minnesota	
Austin, Texas 78712	1	Minneapolis, Minnesota 55455	1
Dr. Keith H. Johnson		Dr. R. P. Van Duyne	
Department of Metallurgy and		Chemistry Department	
Materials Science		Northwestern University	
Massachusetts Institute of Technology	7	Evanston, Illinois 60201	1
Cambridge, Massachusetts 02139	1		
		Dr. S. Sibener	
Dr. J. E. Demuth		Department of Chemistry	
IBM Corporation		James Franck Institute	
Thomas J. Watson Research Center		5640 Ellis Avenue	_
P. O. Box 218		Chicago, Illinois 60637	1
Yorktown Heights, New York 10598	1		
		Dr. M. G. Lagally	
Dr. C. P. Flynn		Department of Metallurgical	
Department of Physics		and Mining Engineering	
University of Illinois		University of Wisconsin	•
Urbana, Illinois 61801	1	Madison, Wisconsin 53706	1

TECHNICAL REPORT DISTRIBUTION LIST, 056

	No. Copies		No. Copies
Dr. Robert Gomer		Dr. K. G. Spears	
Department of Chemistry		Chemistry Department	
James Franck Institute		Northwestern University	
5640 Ellis Avenue		Evanston, Illinois 60201	1
Chicago, Illinois 60637	1	Diameton, IIIInois Vozoi	•
	•	Dr. R. W. Plummer	
Dr. R. G. Wallis		University of Pennsylvania	
Department of Physics		Department of Physics	
University of California, Irvine		Philadelphia, Pennsylvania 19104	1
Irvine, California 92664	1	Introduction I comply that I 1104	•
7100	•	Dr. E. Yeager	
Dr. D. Ramaker		Department of Chemistry	
Chemistry Department		Case Western Reserve University	
George Washington University		Cleveland, Ohio 41106	1
Washington, D.C. 20052	1		•
	•	Professor D. Hercules	
Dr. P. Hansma		University of Pittsburgh	
Physics Department		Chemistry Department	
University of California,		Pittsburgh, Pennsylvania 15260	1
Santa Barbara		11000018m, reimay1vanta 17200	•
Santa Barbara, California 93106	1	Professor N. Winograd	
	_	The Pennsylvania State University	
Dr. J. C. Hemminger		Department of Chemistry	
Chemistry Department		University Park, Pennsylvania 16802	1
University of California, Irvine			•
Irvine, California 92717	1	Professor T. F. George	
•	-	The University of Rochester	
Dr. Martin Fleischmann		Chemistry Department	
Department of Chemistry		Rochester, New York 14627	1
Southampton University			_
Southampton SO9 5NH		Professor Dudley R. Herschbach	
Hampshire, England	1	Harvard College	
• • • • • • • • • • • • • • • • • • • •		Office for Research Contracts	
Dr. G. Rubloff		1350 Massachusetts Avenue	
IBM		Cambridge, Massachusetts 02138	1
Thomas J. Watson Research Center	•		_
P. O. Box 218		Professor Horia Metiu	
Yorktown Heights, New York 10598	1	University of California,	
-		Santa Barbara	
Dr. J. A. Gardner		Chemistry Department	
Department of Physics		Santa Barbara, California 93106	1
Oregon State University			-
Corvallis, Oregon 97331	1	Professor A. Steckl	
- · · · · · · · · · · · · · · · · · · ·	-	Rensselaer Polytechnic Institute	
Dr. G. D. Stein		Department of Electrical and	
Mechanical Engineering Department		Systems Engineering	
Northwestern University		Integrated Circuits Laboratories	
Evanston, Illinois 60201	1	Troy, New York 12181	1
	_		•

TECHNICAL REPORT DISTRIBUTION LIST, 056

Copies Dr. John T. Yates Department of Chemistry University of Pittsburgh Pittsburgh, Pennsylvania 15260 1 Professor G. H. Morrison Department of Chemistry Cornell University Ithaca, New York 14853 1
Department of Chemistry University of Pittsburgh Pittsburgh, Pennsylvania 15260 1 Professor G. H. Morrison Department of Chemistry Cornell University Ithaca, New York 14853 1
Department of Chemistry Cornell University Ithaca, New York 14853 1
Department of Chemistry Cornell University Ithaca, New York 14853 1
Cornell University Ithaca, New York 14853 1
Ithaca, New York 14853
Captain Lee Myers
AFOSR/NC .
Bolling AFB
Washington, D.C. 20332
Dr. David Squire
Army Research Office
P. O. Box 12211
Research Triangle Park, NC 27709 1
Professor Ronald Hoffman
Department of Chemistry
Cornell University
Ithaca, New York 14853

